

of the protons on C'(6) straddle the O(5^I)–C(5^I) bond. The other proton is located then between the methylene protons of C(5^I) and directed towards C(6^{IV}). A similar meshing of the C(6) protons with the O'(5)–C'(5) bond and the methylene protons of C'(5) is necessary to account for the short distances of 3.39 Å and 3.75 Å between C(6^{IV}) and O'(5) and C'(5). If this arrangement of the methyl protons is basically correct, then the approach distance between C'(6) and C(6^I), and C'(6) and C(6^{II}), should be approximately 4.4 Å and 3.7 Å, if the proton van der Waals radius is taken as 1.2 Å. The observed distances are 4.11 Å and 3.62 Å respectively.

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References

- BISHOP, C. T. (1953). *Canad. J. Chem.* **31**, 793.
 BROWN, C. J. (1960). *Acta Cryst.* **13**, 1049.
 BROWN, C. J. (1965). Private communication. To be submitted to *J. Chem. Soc.*
 BROWN, G. M. & LEVY, H. A. (1965). *Science*, **147**, 1038.
- CRUICKSHANK, D. W. J., PILLING, D. E., BUJOSA, A., LOVELL, F. M. & TRUTER, M. R. (1961). In *Computing Methods and the Phase Problem in X-ray Crystal Analysis*, p. 32. Symposia Publications Division, Pergamon Press.
 DAWSON, B. (1960). *Acta Cryst.* **13**, 403.
 FERRIER, W. G. (1963). *Acta Cryst.* **16**, 1023.
 FREEMAN, A. J. (1959). *Acta Cryst.* **12**, 261.
 HOERNI, J. A. & IBERS, J. A. (1954). *Acta Cryst.* **7**, 744.
 HORDVIK, A. (1961). *Acta Chem. Scand.* **15**, 16.
 HOWELLS, E. R., PHILLIPS, D. C. & ROGERS, D. (1950). *Acta Cryst.* **3**, 210.
 JACOBSON, R. A., WUNDERLICH, J. A. & LIPSCOMB, W. N. (1961). *Acta Cryst.* **14**, 598.
 JEFFREY, G. A. & ROSENSTEIN, R. D. (1964). *Advanc. Carb. Chem.* **19**, 7.
 JONES, D. W. (1960). *J. Polymer Sci.* **42**, 173.
 LEMIEUX, R. U. & SHYLUK, W. P. (1953). *Canad. J. Chem.* **31**, 528.
 MANN, J. (1962). *Pure Appl. Chem.* **5**, 91.
 McDONALD, T. R. R. & BEEVERS, C. A. (1952). *Acta Cryst.* **5**, 654.
 MICHELL, A. J. (1965). Private communication.
 PAULING, L. (1960). *The Nature of the Chemical Bond*. 3rd Ed. Birmingham: Kynoch Press.
 REEVES, R. E. (1951). *Advanc. Carb. Chem.* **6**, 107.
 TIMELL, T. E. (1964). *Advanc. Carb. Chem.* **19**, 247.
 WILSON, A. J. C. (1949). *Acta Cryst.* **2**, 318.
 WILSON, A. J. C. (1951). *Research*, **4**, 141.
 YUNDT, A. P. (1951). *Tappi*, **34**, 89.

Acta Cryst. (1966). **21**, 79

Crystal and Molecular Structure of L- α -Glycerylphosphorylcholin

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The crystal structure of L- α -glycerylphosphorylcholin – the basic unit of the lecithins – has been determined and refined to an *R* value of 0.062 by anisotropic least-squares treatment.

Introduction

Our X-ray studies of lipids in the solid state have been undertaken to contribute to the knowledge of the structure of important biological systems with partly ordered lipid molecules such as cell membranes and the myelin sheath of nerves. Phosphatides with zwitterion structure (lecithins, cephalins, sphingomyelins) are important components in such systems. In order to obtain accurate structural information on the polar regions in the lecithins, we have performed a single-crystal analysis of the basic unit, glycerylphosphorylcholin (GPC).

Experimental

Optically active L- α -GPC ($C_8H_{20}O_6NP$) was synthesized according to Baer & Kates (1948). The compound was

purified by four recrystallizations of its $CdCl_2$ complex. GPC was recovered by dissolving the complex in water and passing the solution through an ion exchanger (Tattrie & McArthur, 1958) and finally dried in vacuum over phosphorus pentoxide.

Crystals of GPC were obtained from ethanol (99.5%). They grow in long prisms and are often twinned. As GPC is very hygroscopic the crystals had to be handled in a water-free atmosphere and mounted in glass capillaries for the X-ray work.

The crystals are monoclinic ($P2_1$) with the following cell dimensions (Cu $K\alpha$ radiation): $a = 10.10$, $b = 7.71$, $c = 16.62$ Å, $\beta = 102.7^\circ$.

A reasonable value for the calculated density of the crystals, 1.320 g.cm^{-3} , is obtained if there are four molecules per cell. The Patterson series also conforms with two molecules per asymmetric unit.

Table 1. Fractional atomic coordinates with standard deviations $\times 10^5$ (within brackets) for the heavier atoms of the structure

	<i>x</i>	$\sigma(x)$	<i>y</i>	$\sigma(y)$	<i>z</i>	$\sigma(z)$
P(1)	0.75618	(14)	0.50164	(45)	0.07638	(10)
N(1)	1.17855	(47)	0.46781	(95)	0.12375	(27)
O(1)	0.36688	(34)	0.91207	(84)	0.14787	(22)
O(2)	0.71364	(35)	1.02920	(85)	0.12976	(23)
O(3)	0.68363	(38)	0.68369	(80)	0.05967	(25)
O(4)	0.88628	(37)	0.53917	(84)	0.15116	(21)
O(5)	0.67329	(42)	0.37357	(92)	0.11188	(30)
O(6)	0.80509	(38)	0.45962	(77)	-0.00003	(24)
C(1)	0.49754	(54)	0.99843	(122)	0.16559	(35)
C(2)	0.58712	(50)	0.93828	(109)	0.10709	(32)
C(3)	0.61784	(54)	0.74153	(127)	0.11985	(35)
C(4)	0.98227	(58)	0.66480	(118)	0.13862	(36)
C(5)	1.12703	(53)	0.60826	(106)	0.17234	(30)
C(6)	1.11141	(62)	0.29848	(108)	0.12782	(36)
C(7)	1.16462	(64)	0.52642	(136)	0.03370	(31)
C(8)	1.32862	(54)	0.44543	(134)	0.16279	(45)
P(1')	0.78699	(14)	0.77195	(35)	0.60243	(8)
N(1')	1.20830	(45)	0.84521	(100)	0.63007	(28)
O(1')	0.39722	(39)	0.36879	(83)	0.60716	(23)
O(2')	0.74924	(37)	0.23669	(73)	0.62904	(23)
O(3')	0.71735	(37)	0.58362	(63)	0.59421	(19)
O(4')	0.92882	(36)	0.73692	(74)	0.66865	(21)
O(5')	0.71143	(38)	0.89129	(71)	0.64452	(23)
O(6')	0.81460	(35)	0.80981	(68)	0.52009	(21)
C(1')	0.51099	(35)	0.30496	(103)	0.57736	(33)
C(2')	0.63732	(52)	0.32404	(87)	0.64726	(30)
C(3')	0.68024	(50)	0.50816	(105)	0.66657	(28)
C(4')	1.02850	(56)	0.62383	(103)	0.64619	(38)
C(5')	1.17216	(53)	0.69586	(105)	0.67766	(34)
C(6')	1.12992	(65)	1.00917	(121)	0.63900	(40)
C(7')	1.18714	(83)	0.80890	(144)	0.54095	(40)
C(8')	1.35720	(67)	0.88858	(164)	0.66630	(55)

Table 2. U_{ij} 's together with standard deviations ($\times 10^4$) within brackets

Allowance was made for anisotropic vibration with

$$\exp -2\pi^2(h^2a^{*2} \cdot U_{11} + k^2b^{*2} \cdot U_{22} + l^2c^{*2} \cdot U_{33} + 2kl \cdot b^{*}c^{*} \cdot U_{23} + 2lh^{*}a^{*} \cdot U_{31} + 2hka^{*}b^{*} \cdot U_{12})$$

	U_{11}	U_{22}	U_{33}	U_{23}	U_{13}	U_{12}
P(1)	0.0445 (7)	0.0352 (20)	0.0606 (8)	0.0003 (10)	0.0093 (6)	-0.0026 (10)
N(1)	0.0588 (24)	0.0381 (66)	0.0489 (23)	0.0073 (29)	0.0133 (18)	0.0079 (30)
O(1)	0.0527 (18)	0.0858 (49)	0.0571 (20)	0.0050 (25)	0.0217 (13)	0.0042 (24)
O(2)	0.0566 (19)	0.0299 (43)	0.0790 (20)	-0.0017 (25)	0.0231 (15)	-0.0025 (22)
O(3)	0.0545 (21)	0.0385 (45)	0.0709 (23)	0.0010 (24)	0.0178 (17)	0.0091 (21)
O(4)	0.0535 (19)	0.0504 (44)	0.0494 (18)	-0.0012 (24)	0.0089 (15)	-0.0003 (24)
O(5)	0.0653 (23)	0.0396 (56)	0.1084 (30)	0.0075 (32)	0.0245 (20)	0.0051 (30)
O(6)	0.0653 (21)	0.0435 (50)	0.0646 (21)	-0.0125 (23)	0.0064 (16)	0.0023 (24)
C(1)	0.0584 (29)	0.0641 (71)	0.0719 (31)	0.0015 (40)	0.0234 (22)	0.0174 (37)
C(2)	0.0533 (25)	0.0326 (64)	0.0590 (28)	0.0051 (35)	0.0198 (19)	-0.0060 (31)
C(3)	0.0682 (28)	0.0436 (69)	0.0809 (31)	0.0013 (38)	0.0373 (21)	-0.0073 (36)
C(4)	0.0637 (32)	0.0239 (70)	0.0672 (31)	-0.0137 (37)	0.0174 (24)	-0.0023 (36)
C(5)	0.0593 (30)	0.0302 (67)	0.0430 (25)	-0.0114 (30)	0.0063 (21)	0.0019 (33)
C(6)	0.0828 (34)	0.0187 (66)	0.0652 (31)	0.0005 (36)	0.0245 (24)	0.0011 (36)
C(7)	0.1217 (39)	0.0594 (69)	0.0417 (25)	0.0015 (35)	0.0319 (23)	0.0070 (45)
C(8)	0.0525 (28)	0.0772 (91)	0.0982 (45)	0.0172 (52)	0.0234 (27)	0.0137 (37)
P(1')	0.0482 (6)	0.0314 (16)	0.0451 (6)	0.0009 (8)	0.0128 (4)	-0.0023 (9)
N(1')	0.0610 (25)	0.0605 (62)	0.0567 (24)	0.0165 (29)	0.0185 (19)	-0.0034 (30)
O(1')	0.0620 (22)	0.1248 (49)	0.0582 (20)	-0.0288 (26)	0.0145 (15)	-0.0020 (28)
O(2')	0.0730 (20)	0.0310 (45)	0.0787 (21)	0.0061 (25)	0.0334 (16)	0.0047 (26)
O(3')	0.0659 (21)	0.0357 (40)	0.0429 (16)	-0.0022 (19)	0.0162 (14)	-0.0122 (21)
O(4')	0.0599 (19)	0.0513 (42)	0.0510 (17)	-0.0005 (22)	0.0127 (14)	-0.0086 (24)
O(5')	0.0813 (23)	0.0182 (44)	0.0795 (21)	-0.0024 (24)	0.0320 (16)	-0.0020 (24)
O(6')	0.0545 (19)	0.0674 (45)	0.0587 (19)	0.0086 (23)	0.0114 (15)	-0.0058 (22)
C(1')	0.0664 (28)	0.0511 (64)	0.0599 (28)	-0.0117 (33)	0.0187 (21)	-0.0192 (32)
C(2')	0.0613 (29)	0.0310 (56)	0.0497 (24)	-0.0006 (27)	0.0123 (21)	-0.0036 (28)
C(3')	0.0579 (26)	0.0481 (61)	0.0388 (22)	0.0058 (31)	0.0160 (18)	-0.0000 (34)
C(4')	0.0630 (30)	0.0408 (65)	0.0891 (35)	0.0215 (37)	0.0131 (27)	0.0028 (36)
C(5')	0.0607 (28)	0.0528 (64)	0.0564 (30)	0.0264 (32)	0.0040 (24)	-0.0026 (31)
C(6')	0.0934 (40)	0.0538 (70)	0.0836 (35)	0.0105 (43)	0.0364 (29)	-0.0027 (46)
C(7')	0.1461 (55)	0.0648 (95)	0.0642 (32)	0.0098 (45)	0.0392 (33)	0.0126 (56)
C(8')	0.0556 (35)	0.1569 (104)	0.1489 (56)	0.0621 (69)	0.0189 (36)	-0.0116 (51)

Multiple film Weissenberg photographs were recorded for the 0–4 layers about the *b* axis and the 0 and 2 layers about the *c* axis. The intensities were measured visually and corrected for the Lorentz and polarization factors but not for absorption.

Structure determination

A three-dimensional Patterson series was calculated with coefficients sharpened to correspond to those from point atoms at rest (Abrahamsson & Maslen, 1963). The positions of the two phosphorus atoms were easily derived from the series. They have different *y* coordinates, and no false symmetry is introduced in the electron density series based on the phases of the two atoms. The series in fact showed most of the other atoms in the structure but only the highest peaks were used for the following stage. After four rounds of Fourier refinement all atoms except hydrogen atoms had been located and the *R* value was 0.20.

The structure was further refined by anisotropic least-squares treatment using the full matrix. The progress of the refinement was checked at some stages by calculating difference syntheses. In the first of these, 13 of the hydrogen atoms showed up distinctly and were included in the following structure factor calculations. Each hydrogen atom was assigned an isotropic temperature factor calculated from the anisotropic vibration parameters of the hydrogen-carrying heavier atom. The second difference series calculated at *R*=0.09 gave equally clearly the positions of the remaining hydrogen atoms except H(71'), H(73') and H(81') which, however, were included in the structure factor calculations with their expected coordinates as they belonged to CH₃ groups of which the other hydrogen atoms had been located from the difference maps. After six more least-squares cycles with all atoms of the molecule included the shifts were small (one-third of the standard deviations) and the refinement was stopped. The hydrogen parameters, however, had not been refined. As 289 parameters were varied simultaneously only the right hand sides of the normal equations matrix were calculated in some of the last cycles in order to reduce the computing time and the inverse matrix of an earlier stage was used for forming the shifts.

The final *R* value for the 1597 observed reflexions is 0.062. The scattering curves given in *International*

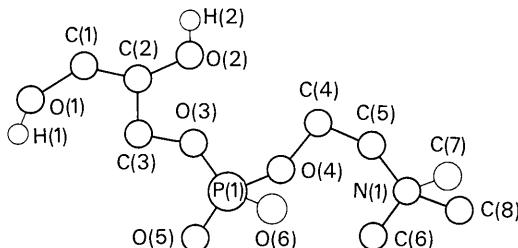


Fig. 1. Numbering of the atoms of the GPC molecule. The two molecules of the asymmetric unit have identical numbering but are differentiated by prime.

Tables for X-ray Crystallography (1962), p. 202, were used. The calculations were performed on the Datasaab D21 computer with the program system developed by Abrahamsson, Aleby, Larsson, Nilsson, Selin & Westerdahl (1965). The weight used for each observation in the least-squares refinement was (Mills & Rollett, 1960)

$$w = \frac{1}{1 + [(|F_o| - 10|F_{\min}|)/4|F_{\min}|]^2}.$$

Discussion

The atomic parameters used in the final calculation of structure factors are given with standard deviations in Tables 1–3. Observed and calculated structure factors are listed in Table 4. The numbering of atoms in the molecule is illustrated in Fig. 1. Primed symbols refer to the second molecule of the asymmetric unit.

The two molecules of the asymmetric unit are shown in Fig. 2. Their conformation and orientation in the unit cell is such that one is roughly related to the other by a *c*-glide plane at *y*=0.13. As the molecule is optically active this relationship is, of course, not valid at

Table 3. Parameters for the hydrogen atoms

	<i>x</i>	<i>y</i>	<i>z</i>	<i>B</i>
H(1)	0.3300	0.8600	0.1000	4.84 Å ²
H(2)	0.7200	1.1550	0.1200	4.07
H(11)	0.4812	0.8620	0.1586	5.02
H(12)	0.5491	0.9694	0.2288	5.02
H(21)	0.5331	0.9603	0.0430	3.89
H(31)	0.6827	0.7191	0.1807	4.81
H(32)	0.5233	0.6684	0.1149	4.81
H(41)	0.9640	0.7839	0.1698	3.72
H(42)	0.9680	0.6884	0.0729	3.72
H(51)	1.1345	0.5587	0.2349	3.12
H(52)	1.1926	0.7200	0.1742	3.12
H(61)	1.1297	0.2558	0.1920	3.34
H(62)	1.0027	0.3120	0.1034	3.34
H(63)	1.1530	0.2038	0.0916	3.34
H(71)	1.0575	0.5516	0.0065	5.21
H(72)	1.2229	0.6438	0.0323	5.21
H(73)	1.2013	0.4240	-0.0007	5.21
H(81)	1.3817	0.5680	0.1593	5.88
H(82)	1.3413	0.4081	0.2273	5.88
H(83)	1.3718	0.3455	0.1299	5.88
H(1')	0.3200	0.4400	0.5700	5.47
H(2')	0.7500	0.1100	0.6300	4.73
H(11')	0.4958	0.1682	0.5607	4.44
H(12')	0.5241	0.3787	0.5239	4.44
H(21')	0.6153	0.2659	0.7027	3.29
H(31')	0.5966	0.5806	0.6820	3.73
H(32')	0.7671	0.5104	0.7187	3.73
H(41')	1.0081	0.6103	0.5796	5.58
H(42')	1.0215	0.4958	0.6740	5.58
H(51')	1.2447	0.5907	0.6749	4.31
H(52')	1.1829	0.7358	0.7414	4.31
H(61')	1.0211	0.9854	0.6172	5.85
H(62')	1.1498	1.0479	0.7038	5.85
H(63')	1.1608	1.1133	0.6027	5.85
H(71')	1.2519	0.6999	0.5317	6.85
H(72')	1.0821	0.7770	0.5160	6.85
H(73')	1.2159	0.9231	0.5097	6.85
H(81')	1.3701	0.9219	0.7311	8.39
H(82')	1.4202	0.7766	0.6600	8.39
H(83')	1.3876	0.9983	0.6329	8.39

Table 4. Observed and calculated structure factors (x100) with phase angles (as fractions of one revolution)

N	K	L	F OBS	F CALC	F I	H	K	L	F OBS	F CALC	F I	H	K	L	F OBS	F CALC	F I	H	K	L	F OBS	F CALC	F I
0	0	0	3753	3499	1.0000	7	0	635	513	1.0000	3	-11	2345	2204	0.6188	-1	-15	621	561	0.2847	-	-	-
0	0	0	20713	19439	0.0000	7	0	635	561	0.0000	3	-14	668	768	0.1705	-	-15	635	561	0.0863	-	-	-
0	0	0	393	372	0.0000	7	0	292	213	0.0000	3	-16	708	1359	0.5443	-	-16	1268	1289	0.7116	-	-	-
0	0	0	100	93	0.0000	7	0	292	213	0.0000	3	-18	428	483	0.4456	-	-18	349	415	0.4550	-	-	-
0	0	0	735	733	0.0000	7	0	477	1201	0.0000	3	-18	1359	1345	0.5443	-	-18	555	570	0.9900	-	-	-
0	0	0	127	123	0.0000	7	0	1561	1484	1.0000	3	-19	401	458	0.3225	-	-19	791	569	0.4451	-	-	-
0	0	0	193	189	0.0000	7	0	1470	1472	0.0000	3	-19	1897	1596	0.1270	-	-19	717	640	0.5695	-	-	-
0	0	0	147	144	0.0000	7	0	1470	1472	0.0000	3	-20	2314	2358	0.4184	-	-20	623	631	0.7530	-	-	-
0	0	0	176	174	0.0000	7	0	453	446	0.0000	3	-20	3637	3411	0.5586	-	-20	1319	1215	0.6559	-	-	-
0	0	0	158	156	0.0000	7	0	453	523	0.0000	3	-20	3447	3447	0.5586	-	-20	603	631	0.7530	-	-	-
0	0	0	185	183	0.0000	7	0	95	253	0.0000	3	-20	2948	2654	0.6089	-	-20	101	101	0.7070	-	-	-
0	0	0	672	683	1.0000	7	0	547	547	0.0000	3	-20	2975	2556	0.5688	-	-20	798	747	0.4456	-	-	-
0	0	0	257	251	0.0000	7	0	470	400	0.0000	3	-20	1374	1251	0.1616	-	-20	1506	1211	0.5144	-	-	-
0	0	0	862	717	0.0000	7	0	1146	1194	1.0000	3	-20	2027	2024	0.4640	-	-20	360	360	0.6090	-	-	-
0	0	0	218	203	0.0000	7	0	1049	987	0.0000	3	-20	969	688	0.6080	-	-20	376	312	0.6132	-	-	-
0	0	0	193	182	0.0000	7	0	1201	1189	0.0000	3	-20	444	419	0.5924	-	-20	255	196	0.7020	-	-	-
0	0	0	1215	1247	0.0000	7	0	1759	1793	1.0000	3	-20	520	610	0.6504	-	-20	675	643	0.5946	-	-	-
0	0	0	1152	1154	0.0000	7	0	427	427	0.0000	3	-20	665	563	0.9129	-	-20	741	621	0.3433	-	-	-
0	0	0	164	162	0.0000	7	0	793	793	0.0000	3	-20	350	350	0.5620	-	-20	681	681	0.7070	-	-	-
0	0	0	538	637	0.0000	7	0	474	600	0.0000	3	-20	259	259	0.6080	-	-20	521	521	0.6588	-	-	-
0	0	0	754	626	0.0000	7	0	180	180	0.0000	3	-20	1030	1342	0.3890	-	-20	652	526	0.4884	-	-	-
0	0	0	248	249	0.0000	7	0	215	210	0.0000	3	-20	761	633	0.4227	-	-20	504	477	0.7171	-	-	-
0	0	0	3768	3718	0.0000	7	0	1045	1072	0.0000	3	-20	1355	1453	0.6045	-	-20	750	492	0.7367	-	-	-
0	0	0	665	665	0.0000	7	0	527	528	0.0000	3	-20	404	403	0.6045	-	-20	454	334	0.2138	-	-	-
0	0	0	1858	1765	0.0000	7	0	523	523	0.0000	3	-20	1537	1558	0.0336	-	-20	769	727	0.1495	-	-	-
0	0	0	426	385	0.0000	7	0	112	115	0.0000	3	-20	862	862	0.1967	-	-20	350	350	0.0942	-	-	-
0	0	0	151	116	0.0000	7	0	218	218	0.0000	3	-20	1582	1773	0.6487	-	-20	473	472	0.4734	-	-	-
0	0	0	745	740	0.0000	7	0	364	614	0.0000	3	-20	689	695	0.1161	-	-20	224	147	0.3335	-	-	-
0	0	0	416	359	0.0000	7	0	565	565	0.0000	3	-20	646	562	0.3179	-	-20	266	236	0.1536	-	-	-
0	0	0	1728	1275	0.0000	7	0	242	242	0.0000	3	-20	714	656	0.0684	-	-20	507	475	0.3764	-	-	-
0	0	0	8610	862	0.0000	7	0	501	506	0.0000	3	-20	394	404	0.2437	-	-20	230	245	0.1161	-	-	-
0	0	0	1503	1507	0.0000	7	0	578	578	0.0000	3	-20	1133	1133	0.5115	-	-20	241	155	0.5167	-	-	-
0	0	0	2953	295	0.0000	7	0	768	765	0.0000	3	-20	3177	3158	0.0718	-	-20	354	292	0.2542	-	-	-
0	0	0	4068	4221	0.0000	7	0	515	874	1.0000	3	-20	1744	1744	0.6077	-	-20	314	274	0.2126	-	-	-
0	0	0	1492	1480	0.0000	7	0	1298	1360	1.0000	3	-20	1609	1673	0.5717	-	-20	473	472	0.4734	-	-	-
0	0	0	1885	1667	0.0000	7	0	915	915	0.0000	3	-20	850	850	0.3165	-	-20	425	372	0.0911	-	-	-
0	0	0	1584	1589	0.0000	7	0	695	695	0.0000	3	-20	535	535	0.2344	-	-20	295	256	0.7656	-	-	-
0	0	0	751	91	0.0000	7	0	455	581	0.0000	3	-20	581	571	0.0893	-	-20	341	353	0.5222	-	-	-
0	0	0	191	122	0.0000	7	0	471	472	0.0000	3	-20	685	572	0.2477	-	-20	151	174	0.7233	-	-	-
0	0	0	590	590	0.0000	7	0	526	526	0.0000	3	-20	488	429	0.4149	-	-20	578	572	0.2211	-	-	-
0	0	0	2584	2584	0.0000	7	0	472	472	0.0000	3	-20	351	223	0.4130	-	-20	425	374	0.3113	-	-	-
0	0	0	2000	2215	0.0000	7	0	811	871	1.0000	3	-20	811	868	0.5919	-	-20	532	532	0.6133	-	-	-
0	0	0	1811	1811	0.0000	7	0	486	486	0.0000	3	-20	1038	1038	0.2454	-	-20	547	481	0.7577	-	-	-
0	0	0	442	509	0.0000	7	0	612	529	0.0000	3	-20	576	603	0.7101	-	-20	205	1763	0.4831	-	-	-
0	0	0	199	359	0.0000	7	0	237	237	0.0000	3	-20	1201	1201	0.5159	-	-20	186	186	0.6123	-	-	-
0	0	0	1397	354	0.0000	7	0	329	415	0.0000	3	-20	1596	1953	0.6958	-	-20	854	612	0.4142	-	-	-
0	0	0	237	354	0.0000	7	0	207	207	0.0000	3	-20	1206	1206	0.5058	-	-20	251	312	0.1124	-	-	-
0	0	0	2682	2682	0.0000	7	0	257	253	0.0000	3	-20	843	775	0.0555	-	-20	251	251	0.5151	-	-	-
0	0	0	6502	6502	0.0000	7	0	349	403	0.0000	3	-20	1205	1205	0.5058	-	-20	255	255	0.0705	-	-	-
0	0	0	2368	2368	0.0000	7	0	2184	2235	0.0000	3	-20	240	245	0.4671	-	-20	265	265	0.2667	-	-	-
0	0	0	2990	2958	0.0000	7	0	121	121	0.0000	3	-20	366	366	0.2462	-	-20	294	278	0.2563	-	-	-
0	0	0	1355	1355	0.0000	7	0	1497	1676	0.0000	3	-20	1207	1207	0.5124	-	-20	301	301	0.1111	-	-	-
0	0	0	1140	1140	0.0000	7	0	2179	2179	0.0000	3	-20	1675	1566	0.3949	-	-20	1665	1392	0.1103	-	-	-
0	0	0	2355	1975	0.0000	7	0	1508	1508	0.0000	3	-20	1771	1771	0.2117	-	-20	160	160	0.6203	-	-	-
0	0	0	1950	1950	0.0000	7	0	2124	2124	0.0000	3	-20	2076	2076	0.2764	-	-20	205	205	0.0729	-	-	-
0	0	0	2357	2357	0.0000	7	0	112	112	0.0000	3	-20	1201	1201	0.5159	-	-20	277	277	0.5111	-	-	-
0	0	0	1766	1766	0.0000	7	0	1588	1588	0.0000	3	-20	1201	1201	0.5159	-	-20	277	277	0.5111	-	-	-
0	0	0	1588	1588	0.0000	7	0	2117	2117	0.0000	3	-20	1201	1201	0.5159	-	-20	277	277	0.5111	-	-	-
0	0	0	1588	1588	0.0000	7	0	1588	1588	0.0000	3	-20	1201	1201	0.5159	-	-20	277	277	0.5111	-	-	-
0	0	0	1588	1588	0.0000	7	0	1588	1588	0.0000	3	-20	1201	1201	0.5159	-	-20	277	277	0.5111	-	-	-
0	0	0	1588	1588	0.0000	7	0	1588	1588	0.0000	3	-20	1201	1201									

Table 4 (*cont.*)

L	POB	PCALC	PI	N	K	L	POB	PCALC	PI	N	K	L	POB	PCALC	PI	N	K	L	POB	PCALC	PI
4	2109	0.2560		10	1	1	POB	PCALC	PI	4	3	1	POB	PCALC	PI	9	1	1	POB	PCALC	PI
4	2616	0.0728		10	1	1	POB	PCALC	PI	5	1	1	POB	PCALC	PI	10	1	1	POB	PCALC	PI
4	1405	0.0000		10	1	1	POB	PCALC	PI	6	1	1	POB	PCALC	PI	11	1	1	POB	PCALC	PI
4	1115	0.0000		10	1	1	POB	PCALC	PI	7	1	1	POB	PCALC	PI	12	1	1	POB	PCALC	PI
4	2032	0.0023		10	1	1	POB	PCALC	PI	8	1	1	POB	PCALC	PI	13	1	1	POB	PCALC	PI
4	2005	0.1364		10	1	1	POB	PCALC	PI	9	1	1	POB	PCALC	PI	14	1	1	POB	PCALC	PI
4	62	0.1316		10	1	1	POB	PCALC	PI	10	1	1	POB	PCALC	PI	15	1	1	POB	PCALC	PI
4	1757	0.1021		10	1	1	POB	PCALC	PI	11	1	1	POB	PCALC	PI	16	1	1	POB	PCALC	PI
4	845	0.0518		10	1	1	POB	PCALC	PI	12	1	1	POB	PCALC	PI	17	1	1	POB	PCALC	PI
4	1000	0.0074		10	1	1	POB	PCALC	PI	13	1	1	POB	PCALC	PI	18	1	1	POB	PCALC	PI
4	1303	0.0000		10	1	1	POB	PCALC	PI	14	1	1	POB	PCALC	PI	19	1	1	POB	PCALC	PI
4	827	0.1092		10	1	1	POB	PCALC	PI	20	1	1	POB	PCALC	PI	21	1	1	POB	PCALC	PI
4	281	0.0817		10	1	1	POB	PCALC	PI	21	1	1	POB	PCALC	PI	22	1	1	POB	PCALC	PI
4	339	0.0466		10	1	1	POB	PCALC	PI	22	1	1	POB	PCALC	PI	23	1	1	POB	PCALC	PI
4	843	0.0617		10	1	1	POB	PCALC	PI	23	1	1	POB	PCALC	PI	24	1	1	POB	PCALC	PI
4	1143	0.0051		10	1	1	POB	PCALC	PI	24	1	1	POB	PCALC	PI	25	1	1	POB	PCALC	PI
4	1004	0.0000		10	1	1	POB	PCALC	PI	25	1	1	POB	PCALC	PI	26	1	1	POB	PCALC	PI
4	2085	0.0850		10	1	1	POB	PCALC	PI	26	1	1	POB	PCALC	PI	27	1	1	POB	PCALC	PI
4	1472	0.2166		10	1	1	POB	PCALC	PI	27	1	1	POB	PCALC	PI	28	1	1	POB	PCALC	PI
4	1564	0.1107		10	1	1	POB	PCALC	PI	28	1	1	POB	PCALC	PI	29	1	1	POB	PCALC	PI
4	881	0.0765		10	1	1	POB	PCALC	PI	29	1	1	POB	PCALC	PI	30	1	1	POB	PCALC	PI
4	1556	0.1346		10	1	1	POB	PCALC	PI	30	1	1	POB	PCALC	PI	31	1	1	POB	PCALC	PI
4	1118	0.0732		10	1	1	POB	PCALC	PI	31	1	1	POB	PCALC	PI	32	1	1	POB	PCALC	PI
4	1255	0.0000		10	1	1	POB	PCALC	PI	32	1	1	POB	PCALC	PI	33	1	1	POB	PCALC	PI
4	691	0.1653		10	1	1	POB	PCALC	PI	33	1	1	POB	PCALC	PI	34	1	1	POB	PCALC	PI
4	553	0.0357		10	1	1	POB	PCALC	PI	34	1	1	POB	PCALC	PI	35	1	1	POB	PCALC	PI
4	273	0.0756		10	1	1	POB	PCALC	PI	35	1	1	POB	PCALC	PI	36	1	1	POB	PCALC	PI
4	401	0.1577		10	1	1	POB	PCALC	PI	36	1	1	POB	PCALC	PI	37	1	1	POB	PCALC	PI
4	609	0.1231		10	1	1	POB	PCALC	PI	37	1	1	POB	PCALC	PI	38	1	1	POB	PCALC	PI
4	543	0.2392		10	1	1	POB	PCALC	PI	38	1	1	POB	PCALC	PI	39	1	1	POB	PCALC	PI
4	622	0.0651		10	1	1	POB	PCALC	PI	39	1	1	POB	PCALC	PI	40	1	1	POB	PCALC	PI
4	1764	0.1446		10	1	1	POB	PCALC	PI	40	1	1	POB	PCALC	PI	41	1	1	POB	PCALC	PI
4	713	0.2467		10	1	1	POB	PCALC	PI	41	1	1	POB	PCALC	PI	42	1	1	POB	PCALC	PI
4	1095	0.0000		10	1	1	POB	PCALC	PI	42	1	1	POB	PCALC	PI	43	1	1	POB	PCALC	PI
4	1508	0.0507		10	1	1	POB	PCALC	PI	43	1	1	POB	PCALC	PI	44	1	1	POB	PCALC	PI
4	2086	0.0000		10	1	1	POB	PCALC	PI	44	1	1	POB	PCALC	PI	45	1	1	POB	PCALC	PI
4	301	0.0586		10	1	1	POB	PCALC	PI	45	1	1	POB	PCALC	PI	46	1	1	POB	PCALC	PI
4	877	0.0461		10	1	1	POB	PCALC	PI	46	1	1	POB	PCALC	PI	47	1	1	POB	PCALC	PI
4	1379	0.0636		10	1	1	POB	PCALC	PI	47	1	1	POB	PCALC	PI	48	1	1	POB	PCALC	PI
4	2763	0.0741		10	1	1	POB	PCALC	PI	48	1	1	POB	PCALC	PI	49	1	1	POB	PCALC	PI
4	2033	0.0741		10	1	1	POB	PCALC	PI	49	1	1	POB	PCALC	PI	50	1	1	POB	PCALC	PI
4	744	0.0265		10	1	1	POB	PCALC	PI	50	1	1	POB	PCALC	PI	51	1	1	POB	PCALC	PI
4	701	0.2580		10	1	1	POB	PCALC	PI	51	1	1	POB	PCALC	PI	52	1	1	POB	PCALC	PI
4	744	0.0449		10	1	1	POB	PCALC	PI	52	1	1	POB	PCALC	PI	53	1	1	POB	PCALC	PI
4	626	0.0000		10	1	1	POB	PCALC	PI	53	1	1	POB	PCALC	PI	54	1	1	POB	PCALC	PI
4	159	0.0000		10	1	1	POB	PCALC	PI	54	1	1	POB	PCALC	PI	55	1	1	POB	PCALC	PI
4	137	0.0416		10	1	1	POB	PCALC	PI	55	1	1	POB	PCALC	PI	56	1	1	POB	PCALC	PI
4	1515	0.0263		10	1	1	POB	PCALC	PI	56	1	1	POB	PCALC	PI	57	1	1	POB	PCALC	PI
4	147	0.1506		10	1	1	POB	PCALC	PI	57	1	1	POB	PCALC	PI	58	1	1	POB	PCALC	PI
4	1507	0.0887		10	1	1	POB	PCALC	PI	58	1	1	POB	PCALC	PI	59	1	1	POB	PCALC	PI
4	2150	0.0000		10	1	1	POB	PCALC	PI	59	1	1	POB	PCALC	PI	60	1	1	POB	PCALC	PI
4	445	0.0506		10	1	1	POB	PCALC	PI	60	1	1	POB	PCALC	PI	61	1	1	POB	PCALC	PI
4	2073	0.0000		10	1	1	POB	PCALC	PI	61	1	1	POB	PCALC	PI	62	1	1	POB	PCALC	PI
4	351	0.0742		10	1	1	POB	PCALC	PI	62	1	1	POB	PCALC	PI	63	1	1	POB	PCALC	PI
4	544	0.0813		10	1	1	POB	PCALC	PI	63	1	1	POB	PCALC	PI	64	1	1	POB	PCALC	PI
4	2417	0.2615		10	1	1	POB	PCALC	PI	64	1	1	POB	PCALC	PI	65	1	1	POB	PCALC	PI
4	1939	0.0000		10	1	1	POB	PCALC	PI	65	1	1	POB	PCALC	PI	66	1	1	POB	PCALC	PI
4	1122	0.2495		10	1	1	POB	PCALC	PI	66	1	1	POB	PCALC	PI	67	1	1	POB	PCALC	PI
4	1256	0.0000		10	1	1	POB	PCALC	PI	67	1	1	POB	PCALC	PI	68	1	1	POB	PCALC	PI
4	1066	0.0000		10	1	1	POB	PCALC	PI	68	1	1	POB	PCALC	PI	69	1	1	POB	PCALC	PI
4	2250	0.0532		10	1	1	POB	PCALC	PI	69	1	1	POB	PCALC	PI	70	1	1	POB	PCALC	PI
4	1571	0.0193		10	1	1	POB	PCALC	PI	70	1	1	POB	PCALC	PI	71	1	1	POB	PCALC	PI
4	661	0.1915		10	1	1	POB	PCALC	PI	71	1	1	POB	PCALC	PI	72	1	1	POB	PCALC	PI
4	421	0.0577		10	1	1	POB	PCALC	PI	72	1	1	POB	PCALC	PI	73	1	1	POB	PCALC	PI
4	875	0.0706		10	1	1	POB	PCALC	PI	73	1	1	POB	PCALC	PI	74	1	1	POB	PCALC	PI
4	243	0.0000		10	1	1	POB	PCALC	PI	74	1	1	POB	PCALC	PI	75	1	1	POB	PCALC	PI
4	1145	0.0000		10	1	1	POB	PCALC	PI	75	1	1	POB	PCALC	PI	76	1	1	POB	PCALC	PI
4	1745	0.0477		10	1	1	POB	PCALC	PI	76	1	1	POB	PCALC	PI	77	1	1	POB	PCALC	PI
4	310	0.0465		10	1	1	POB	PCALC	PI	77	1	1	POB	PCALC	PI	78	1	1	POB	PCALC	PI
4	475	0.0297		10	1	1	POB	PCALC	PI	78	1	1	POB	PCALC	PI	79	1	1	POB	PCALC	PI
4	316	0.0396		10	1	1	POB	PCALC	PI	79	1	1	POB	PCALC	PI	80	1	1	POB	PCALC	PI
4	2145	0.0000		10	1	1	POB	PCALC	PI	80	1	1	POB	PCALC	PI	81	1	1	POB	PCALC	PI
4	1146	0.0000		10	1	1	POB	PCALC	PI	81	1	1	POB	PCALC	PI	82	1	1	POB	PCALC	PI
4	1245	0.0000		10	1	1	POB	PCALC	PI	82	1	1	POB	PCALC	PI	83	1	1	POB	PCALC	PI
4	1147	0.0000		10	1	1	POB	PCALC	PI	83	1	1	POB	PCALC	PI	84	1	1	POB	PCALC	PI
4	1246	0.0000		10	1	1	POB	PCALC	PI	84	1	1	POB	PCALC	PI	85	1	1	POB	PCALC	PI
4	1148	0.0000		10	1	1	POB	PCALC	PI	85	1	1	POB	PCALC	PI	86	1	1	POB	PCALC	PI
4	1247	0.0000		10	1	1	POB	PCAL													

Table 4 (cont.)

H	K	L	POBS	PCALC	PI	H	K	L	POBS	PCALC	PI	H	K	L	POBS	PCALC	PI	H	K	L	POBS	PCALC	PI
6	4	-9	766	810	0.1843	7	4	-8	260	257	0.1440	8	4	-15	360	378	0.3598	10	4	1	667	625	0.6512
6	4	-11	765	810	0.2681	7	4	-5	541	570	0.0761	8	4	-12	243	241	0.2549	10	4	2	429	389	0.1992
6	4	-12	765	713	0.5616	7	4	-11	623	623	0.6427	8	4	-11	306	377	0.2617	10	4	3	518	476	0.7762
6	4	-13	765	201	0.3192	7	4	-10	510	510	0.5105	8	4	-10	1016	973	0.3605	10	4	4	446	376	0.4466
6	4	-15	148	248	0.5605	7	4	-8	170	170	0.5105	8	4	-8	125	133	0.3605	10	4	5	276	228	0.0909
7	4	-17	181	248	0.5603	7	4	-6	262	262	0.5105	8	4	-6	265	198	0.3598	10	4	6	222	231	0.7557
7	4	-18	1066	1159	0.5443	7	4	-4	525	525	0.0601	8	4	-4	303	331	0.3605	10	4	7	304	265	0.6333
7	4	-19	1789	818	0.2421	7	4	-2	1814	1814	0.1253	8	4	-2	1095	1009	0.0531	10	4	8	465	360	0.4001
7	4	-20	1281	1860	0.2559	7	4	0	536	536	0.6530	8	4	0	116	116	0.0531	10	4	9	376	358	0.3591
7	4	-21	1281	1861	0.2559	7	4	-1	536	536	0.6530	8	4	-1	601	545	0.0534	10	4	-6	522	494	0.0551
7	4	-22	443	465	0.0658	7	4	-3	470	470	0.6568	8	4	-3	391	391	0.3598	10	4	-7	511	511	0.3123
7	4	-23	455	455	0.0658	7	4	-5	270	270	0.6568	8	4	-5	266	247	0.4342	10	4	-10	532	563	0.3672
7	4	-24	1056	1056	0.2041	7	4	-7	623	623	0.6568	8	4	-7	1055	1055	0.3598	10	4	-12	221	221	0.3476
7	4	-25	287	288	0.5810	7	4	-9	623	623	0.5311	8	4	-9	1241	1241	0.3598	10	4	-13	361	376	0.2603
7	4	-26	525	525	0.7496	7	4	-11	444	444	0.0480	8	4	-11	863	845	0.6132	10	4	-14	242	186	0.3136
7	4	-27	1727	1773	0.2406	7	4	-13	1146	1221	0.0486	8	4	-13	681	612	0.7024	10	4	-15	242	261	0.1977
7	4	-28	875	881	0.5913	7	4	-14	690	730	0.4662	8	4	-14	217	211	0.2245	10	4	-16	242	242	0.1975
7	4	-29	1076	1081	0.5913	7	4	-16	1009	1034	0.5801	8	4	-16	607	607	0.3598	10	4	-17	205	205	0.6145
7	4	-30	991	1046	0.4663	7	4	-18	520	520	0.3597	8	4	-18	806	716	0.3597	10	4	-18	263	259	0.3246
7	4	-31	1216	1216	0.2440	7	4	-20	303	303	0.3597	8	4	-20	266	266	0.3597	10	4	-20	221	221	0.3476

the asymmetric glycerol carbon atom. The differences in conformation are illustrated in Fig. 3 showing projections down the C(2)-C(3) and C(2')-C(3') bonds. The indicated dihedral angles $\varphi_{O(2)O(3)}$ and $\varphi_{O(2)C(1)}$ for one molecule are 71° and 179° respectively. The corresponding angles for the other GPC molecule of the asymmetric unit are 61° and 63° . In the first case a

gauche-trans conformation has thus been adopted and a gauche-gauche conformation in the latter case. These arrangements also make possible an effective hydrogen bond system similar for both independent molecules.

As observed in 2-aminoethanol phosphate (Kraut, 1961) a gauche conformation is adopted about the bond between the two ethyl carbon atoms of the nitrogen

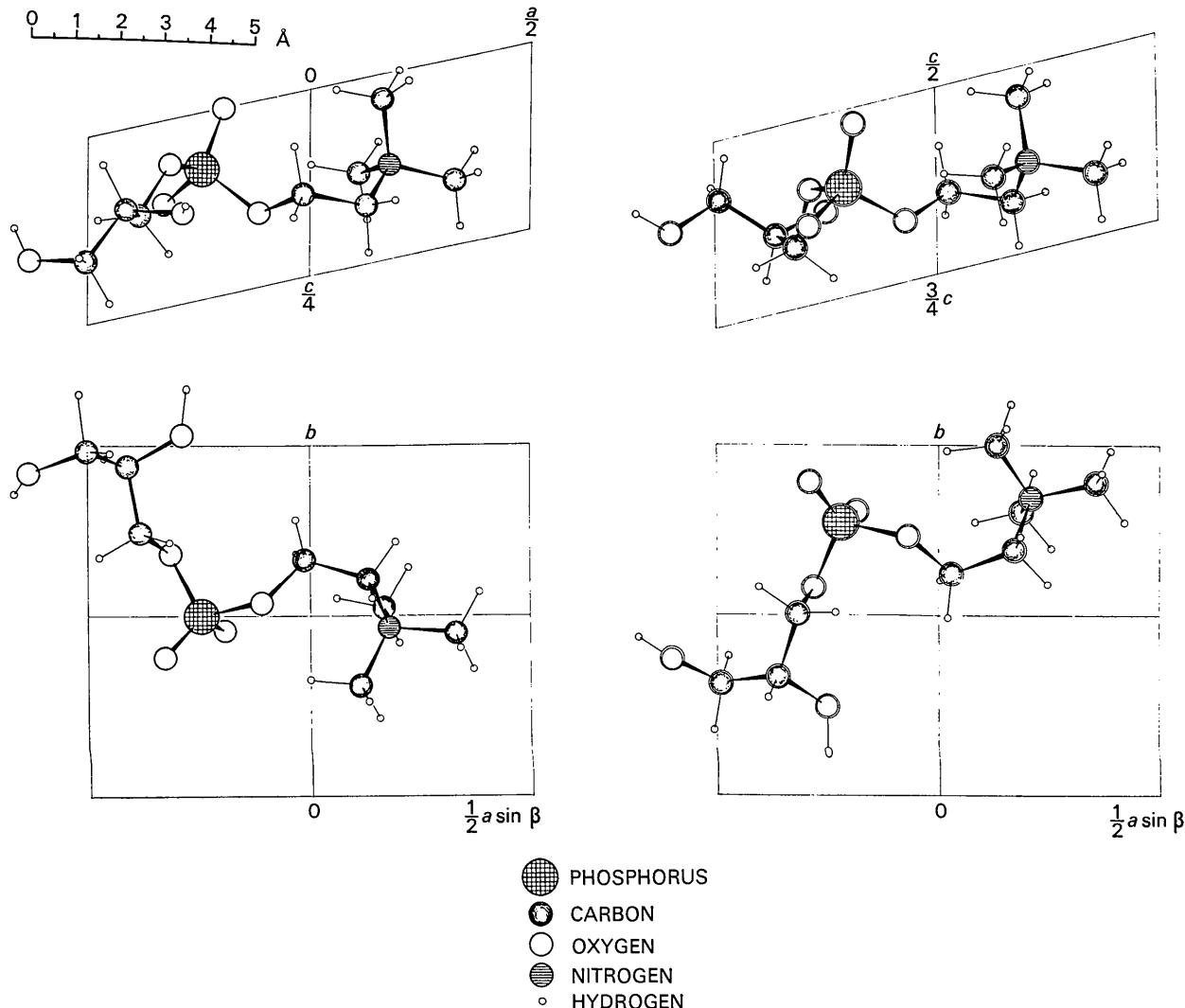


Fig. 2. Spatial drawings of the two GPC molecules of the asymmetric unit. All atoms except hydrogen atoms of one of the two molecules are marked with double contours.

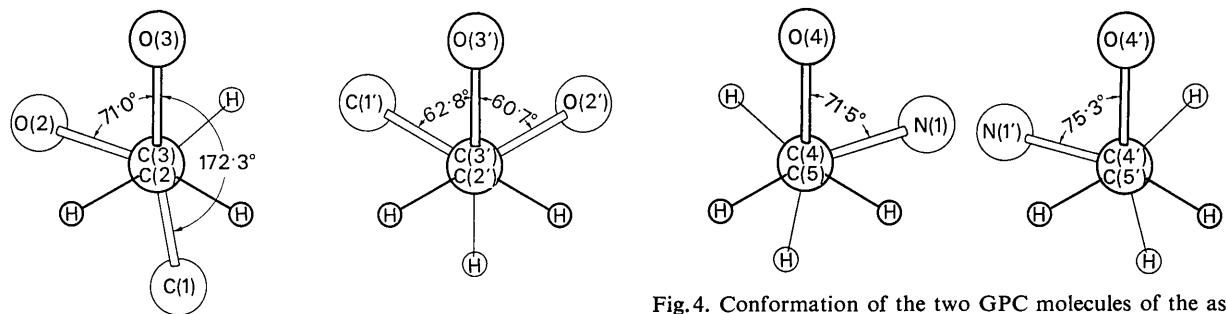


Fig. 3. Conformation of the independent GPC molecules as seen along the C(3)-C(2) and C(3')-C(2') bonds. C(2) and C(2') represent the asymmetric carbon atoms of the molecules.

Fig. 4. Conformation of the two GPC molecules of the asymmetric unit as seen along the C(4)-C(5) and C(4')-C(5') bonds.

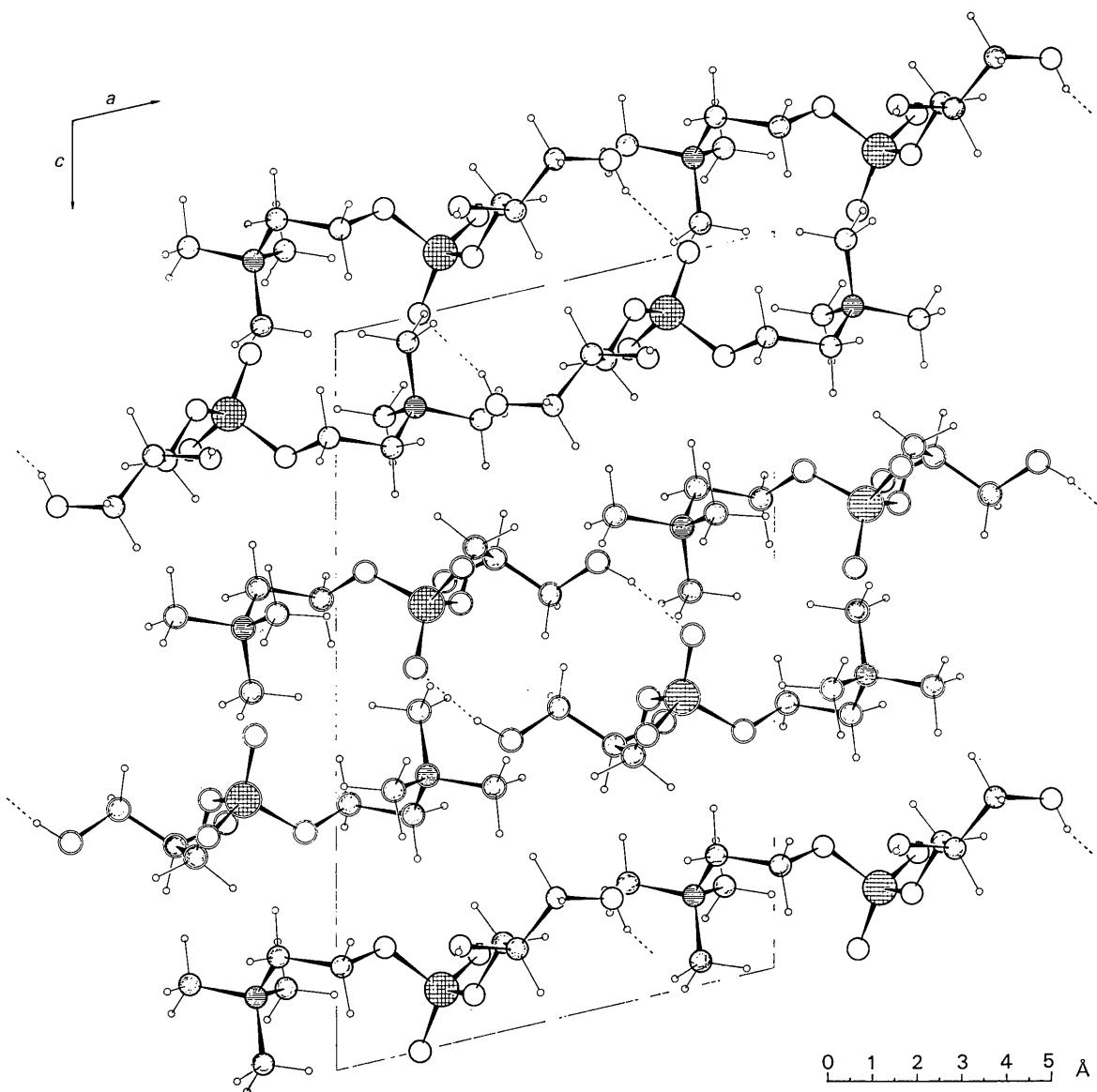


Fig. 5. Illustration of the molecular packing of GPC as seen along the *b* axis. Dashed lines indicate hydrogen bonds. For key to atoms see Fig. 2.

base. The dihedral angles $\varphi_{O(4)N(1)}$ and $\varphi_{O(4')N(1')}$ as seen along the C(4)-C(5) and C(4')-C(5') bonds (Fig. 4) are 72° and 75° respectively.

Bond distances and angles for the two independent molecules of the asymmetric unit are given in Tables 5 and 6. Standard deviations in distances and angles are also listed. They were calculated according to Ahmed & Cruickshank (1953) and Darlow (1960). The maximum deviations from the mean values of identical bonds of the two molecules of the asymmetric unit are, except in a few cases, smaller than 3σ . The standard deviations are, however, slightly underestimated as they have been calculated from the diagonal elements of the inverse least-squares matrix.

The distances in the phosphate group agree well with the recent analyses of adenosine-5'-phosphate (Kraut & Jensen, 1963) and di-*p*-chlorophenyl hydrogen phosphate (DPDPHP) (Calleri & Speakman, 1964). In

their paper, Calleri & Speakman surveyed the P-O distances in sixteen earlier accurate structure determinations and calculated that the mean value of the sum of the four P-O distances of the phosphate group was 6.177 ± 0.030 Å. This sum should also be almost constant according to Cruickshank (1961). The value for adenosine 5'-phosphate is 6.185 Å, for DPCPHP 6.151 Å, and for the two GPC molecules 6.196 Å and 6.184 Å.

It is often stated in the literature that each molecule of GPC is associated with a molecule of water. This is obviously not the case in this structure. The zwitterion character of the molecule is evident from the two equivalent short distances between the phosphorus atom and the unsubstituted oxygen atoms [O(5) and O(6)]. The corresponding O-P-O angle is larger in GPC (mean value 121.6°) than adenosine 5'-phosphate (118°). The charge distribution is also different

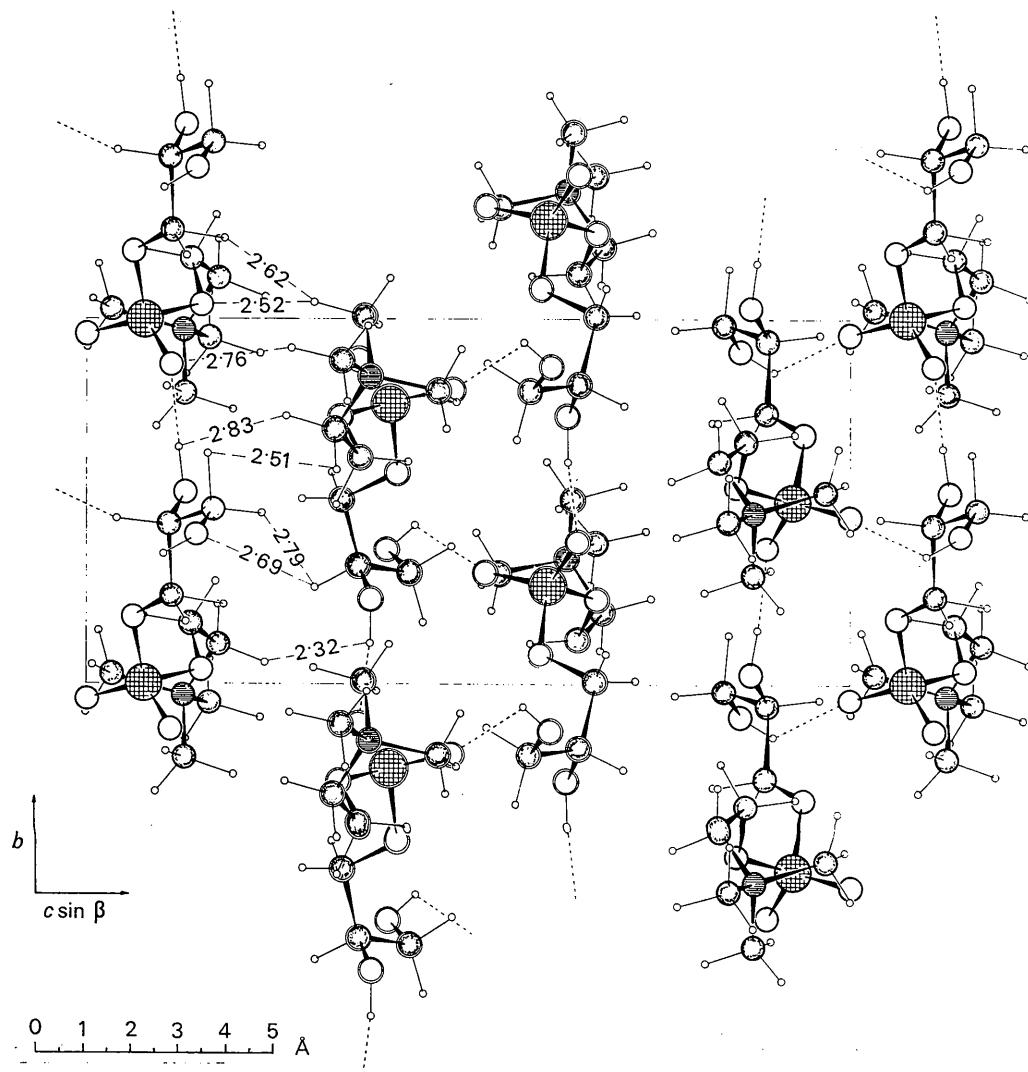


Fig. 6. Molecular packing of GPC as seen along the α axis. The hydrogen bonds forming the double layers are indicated as well as some of the shorter contacts between these layers. For key to atoms see Fig. 2.

Table 5. Bond distances with standard deviations for the two independent molecules of the asymmetric unit

Bond	Molecule 1		Molecule 1'	
	Length	σ	Length	σ
P(1)–O(3)	1.580 Å	0.0065	1.606 Å	0.0053
–O(4)	1.624	0.0037	1.626	0.0035
–O(5)	1.497	0.0065	1.467	0.0052
–O(6)	1.496	0.0047	1.484	0.0041
N(1)–C(5)	1.510	0.0095	1.487	0.0100
–C(6)	1.480	0.0105	1.516	0.0111
–C(7)	1.540	0.0076	1.476	0.0083
–C(8)	1.522	0.0069	1.529	0.0082
O(1)–C(1)	1.449	0.0077	1.435	0.0077
O(2)–C(2)	1.433	0.0073	1.405	0.0073
O(3)–C(3)	1.390	0.0082	1.457	0.0068
O(4)–C(4)	1.418	0.0093	1.442	0.0083
C(1)–C(2)	1.538	0.0089	1.532	0.0067
C(2)–C(3)	1.553	0.0127	1.498	0.0102
C(4)–C(5)	1.510	0.0080	1.534	0.0080

Table 6. Bond angles with standard deviations in degrees for the two independent molecules of the asymmetric unit

Angle	Molecule 1		Molecule 1'	
	θ	$\sigma(\theta)$	θ	$\sigma(\theta)$
O(3)–P(1)–O(4)	104.02°	0.32°	102.22°	0.27°
–O(5)	112.17	0.30	110.09	0.26
–O(6)	105.68	0.30	105.48	0.25
O(4)–P(1)–O(5)	104.15	0.27	104.75	0.23
–O(6)	109.01	0.22	109.91	0.21
O(5)–P(1)–O(6)	120.62	0.38	122.62	0.31
C(5)–N(1)–C(6)	113.14	0.50	113.52	0.51
–C(7)	109.92	0.60	112.86	0.66
–C(8)	106.34	0.50	107.16	0.56
C(6)–N(1)–C(7)	111.01	0.52	107.03	0.57
–C(8)	107.70	0.61	105.58	0.67
C(7)–N(1)–C(8)	108.50	0.49	110.46	0.58
P(1)–O(3)–C(3)	115.80	0.49	118.56	0.37
P(1)–O(4)–C(4)	118.27	0.37	119.13	0.34
O(1)–C(1)–C(2)	111.65	0.56	107.58	0.45
O(2)–C(2)–C(1)	107.50	0.53	111.47	0.46
–C(3)	107.57	0.47	106.96	0.46
C(1)–C(2)–C(3)	109.62	0.58	114.01	0.51
O(3)–C(3)–C(2)	109.13	0.58	108.44	0.45
O(4)–C(4)–C(5)	112.69	0.65	110.78	0.57
N(1)–C(5)–C(4)	115.33	0.47	115.04	0.47

in the phosphate groups of the two compounds as one of the remaining oxygen atoms in adenosine 5'-phosphate is bonded to a hydrogen atom.

The molecular packing is illustrated in Figs. 5 and 6. There are no intramolecular hydrogen bonds. In both independent molecules the glycerol oxygen atom O(1) forms a hydrogen bond of 2.70 Å with O(6) of the phosphate group of a symmetry-related molecule. The other hydroxyl oxygen atom O(2) of the two glycerol residues takes part in a hydrogen bond (2.70 Å) almost

parallel to the *b* axis with O(5) of a *b*-translated molecule. These hydrogen bonds link together equivalent molecules into infinite spirals in the **b** direction. This also conforms with **b** being the needle direction. There are no hydrogen bonds between the two molecules constituting the asymmetric unit.

The structure can also be described as being built up of bimolecular layers parallel to the *ab* plane of one type of molecule alternating with double layers of molecules of the other type. These layers show fairly smooth boundary surfaces towards each other. Some of the shorter interlayer distances are indicated in Fig. 6.

The positively charged nitrogen atom is roughly tetrahedrally surrounded by oxygen atoms. The two shortest N–O distances are between N(1) and the negatively charged oxygen atoms O(6) and O(5') (3.81 Å and 3.88 Å). These atoms lie very close to triad axes of the tetramethylammonium tetrahedron. One hydroxyl oxygen atom O(1) is also fairly close to N(1) (3.90 Å), whereas the fourth oxygen atom O(6) of a symmetry-related molecule is further away (4.44 Å). The packing around N(1') seems less effective as the corresponding shortest N–O distances are about 0.3 Å longer here.

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References

- ABRAHAMSSON, S. & MASLEN, E. N. (1963). *Z. Kristallogr.* **118**, 1.
- ABRAHAMSSON, S., ALEBY, S., LARSSON, K., NILSSON, B., SELIN, K. & WESTERDAHL, A. (1965). *Acta Chem. Scand.* **19**, 758.
- AHMED, F. R. & CRUICKSHANK, D. W. J. (1953). *Acta Cryst.* **6**, 385.
- BAER, E. & KATES, M. (1948). *J. Amer. Chem. Soc.* **70**, 1394.
- CALLERI, M. & SPEAKMAN, J. C. (1964). *Acta Cryst.* **17**, 1097.
- CRUICKSHANK, D. W. J. (1961). *J. Chem. Soc.* p. 5486.
- DARLOW, S. F. (1960). *Acta Cryst.* **13**, 683.
- KRAUT, J. (1961). *Acta Cryst.* **14**, 1146.
- KRAUT, J. & JENSEN, L. H. (1963). *Acta Cryst.* **16**, 79.
- MILLS, O. S. & ROLLETT, J. S. (1960). In *Computing methods and the Phase Problem in X-Ray Crystal Analysis*, p. 107. London: Pergamon Press.
- TATTRIE, N. H. & McARTHUR, C. S. (1958). *Biochem. Prep.* **6**, 16.